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162 CHAPTER 6: THERMOCHEMISTRY To Convert The Answer To Joules, We Write: $101.3 \text{ J} / 0.18 \text{ L Atm} = 1 \text{ L Atm} = - \cdot \times = \cdot \text{ W} - 18 \text{ J}$ 6.17 An Expansion Implies An Increase In Volume, Therefore W Must Be -325 J (see The Defining Equation For Pressure-volume Work.) If The System Absorbs Heat, Q Must Be $+127 \text{ J}$. The Change In Energy (internal 4th, 2024

Thermochemistry Annotated Chapter

6.6 Enthalpy: The Heat Evolved In A Chemical Reaction At Constant Pressure 210 6.7 Constant-Pressure Calorimetry: Measuring ΔH 214 6.8 Relationships Involving ΔH 216 6.9 Determining Enthalpies Of Reaction From Standard Enthalpies Of Formation 218 6.10 Energy Use 6th, 2024

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