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Chapter 2 Electroanalytical Methods- Potentiometry
Standard Potential For Cell, E₀ Cell, Follows Nernst Equation: $E_{\text{Cell}} = (RT/nF) \ln K$ R Is Gas Law Constant, T Is Temperature, K Is Equilibrium Constant, F Is The Faraday, And N Is Number Of Equivalents Of Electricity
Electrode Potential Is Sum Of Two Half-reactions. Cannot Determine Potential Of A Single Electrode, Measure Differences In Potential.
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The Net Cell Reaction That Occurs In The Cell Shown In Fig. 22-1 Is The Sum Of The Two Half-cell Reactions
 $Zn(s) + Cu^{2+} \rightarrow Zn^{2+} + Cu(s)$ Cells, That Are Open Feb 1th, 2024.

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 $E_{\text{Cell}} = E_{\text{Ind}} - E_{\text{Ref}} (+ E_{\text{J}})$ Nernst Equation
 $RT \ln \frac{[Ox]}{[Red]} = E - E^{\circ} - \frac{RT}{nF} \ln \frac{[Red]}{[Ox]}$ Where R = Gas Constant T = Absolute Temperature N = Number Of Electrons In Reaction F = Faraday's Constant E = Potential E° = Standard Potential [Red] = Molar Concentration Of Reduced Form Of Spec

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