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$\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{H}^+} a_{\text{O}_2}}{a_{\text{H}_2}}$  (2.1.36) The Free Energy Of The Cell Reaction Is Given By,  $\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{R}} a_{\text{A}}}{a_{\text{N}} a_{\text{H}^+} a_{\text{O}_2}}$  (2.1.38) Which Since  $\Delta G = -nF E_{\text{rxn}}$  And  $a_{\text{H}^+} = a_{\text{H}_2} = 1$  Becomes,  $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{R}} a_{\text{O}_2}}{a_{\text{A}}}$ . (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can 14th, 2024

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