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$\text{N}2 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{N}2\text{H}4$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{N}2} a_{\text{H}^+}^2}{a_{\text{N}2\text{H}4}}$ (2.1.38) Which Since $\Delta G = -nFE_{\text{rxn}}$ And $a_{\text{H}^+} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{N}2}}{a_{\text{N}2\text{H}4}}$ (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can 8th, 2024

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