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### **Thermodynamic Variables Thermodynamic System**

Thermodynamic Properties August 26, 2010 ME 370 – Thermodynamics 1 Unit One – Properties Of Pure ... Gases” Such As Air 11 Ideal Gas Calculation II • A Volume Of 20 Ft<sup>3</sup> At A Pressure Of ... – R-134a Tables In English Units 4th, 2024

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AlF<sub>4</sub>Na Sodium Tetrafluoroaluminate -1869.0 -1827.5  
345.7 105.9 AlH Aluminum Hydride 259.2 231.2 187.9  
29.4 AlH<sub>3</sub> Aluminum Hydride -46.0 30.0 40.2 AlH<sub>4</sub>K  
Potassium Aluminum Hydride -183.7 AlH<sub>4</sub>Li Lithium  
Aluminum Hydride -116.3 -44.7 78.7 83.2 AlH<sub>4</sub>Na  
Sodium Aluminum Hydride -15.5 Al Aluminum  
Monoiodide 65.5 36.0 4th, 2024

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## **STANDARD THERMODYNAMIC PROPERTIES OF CHEMICAL ...**

NIST Chemistry Webbook, Molecular Formula Crystal  
Liquid Gas Name F H ° KJ/mol F G ° KJ/mol S ° J/mol K C  
P J/mol K F H ° KJ/mol F G ° KJ/mol S ° J/mol K C P J/mol

K F H ° KJ/mol F G ° KJ/mol S ° J/mol K C P J/mol K  
 Substances Not Containing Carbon: Ac Actinium 0.0  
 56.5 27.2 406.0 366.0 188.1 20.8 Ag Silver 0.0 42.6 ...  
 9th, 2024

## **Thermodynamic Stability: Free Energy And Chemical Equilib-** ...

1. Examples of Free Energy Calculations Free Energy Calculations Are Carried Out In Much The Same As Enthalpy calculations. There Are Tables Of Standard Free Energies Of Formation Of Compounds. Elements In Their Standard States Are assigned Zero as their Gibbs Free Energy Of Formation. Consider The Following Chemical Reaction:  $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl}(\text{g})$ . 2th, 2024

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Standard Enthalpy Of Formation And Energy Of HOMO. The Second PC Is Strongly Influenced By Energy Of LUMO, While Substitution Pattern Parameters, Number Of Chlorine Atoms At  $\alpha$ -positions And Vicinal (adjacent) Carbon Atoms 8th, 2024

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Other than The standard state, Value of  $\Delta H_f^\circ$  represents the apparent Enthalpy Of the Reaction Of Formation Of the Solution From The Elements Comprising The solute, ... 2th, 2024

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1 Chemical Reactions 1.1 Rate Of Reaction And Dependence On Temperature We Will Once Again Look At The Formation Of Ammonia ( $\text{NH}_3$ ) From Nitrogen And Hydrogen (see Section Chemical Equilibrium Of The Thermodynamics Chapter). This Reaction Follows The Equation:  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  (1)  $H_0 = 92 \text{ KJ Mol}^{-1}$   $S_0 = 192 \text{ J Mol}^{-1} \text{ K}^{-1}$  1th, 2024

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## **Reaction ...**

Grading: Exam 1 15% Exam 2 15% Exam 3 15% Final Exam 20% Research Paper 10% Homework 20% Class Participation 5% Last Year, The Following Grading Scale Was Used:  $A \geq 92$   $A/B \geq 89$   $B \geq 82$   $B/C \geq 80$   $C \geq 71$   $C/D \geq 69$   $D \geq 61$   $F \leq 60$  If Necessary, The Grading Scale May Be Curved Downward. 6th, 2024

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