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 $\% \text{ Yield} = \text{Actual Yield} \times 100 \%$
 $\text{Theoretical Yield Moles Of } \dots \Delta G^\circ = -RT \ln K$
Standard Emf Of An Electrochemical Cell: $E^\circ \text{ Cell} = E^\circ \text{ Ox} - E^\circ \text{ Red} = E^\circ \text{ Cathode} - E^\circ \text{ Anode}$
Standard Free Energy Change: $\Delta G^\circ = -nFE^\circ \text{ Cell}$ Where F Is The Faraday Constant
Relationship Of The Standard Emf Of The Cell To The Equilibrium Constant: $E^\circ \text{ Cell} = \frac{RT}{nF} \ln K$
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