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How To Do Lab Report [Exp 004] Rates Of Reaction For Iodine Clock Reaction
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$= k[D]_1^x[B]_1^y$ $k[D]_2^x[B]_2^y$ Equation 5 In This Equation k Cancels Out. $[D]_1 = [D]_2$ Because We Are Starting Off With The Same Initial Concentration Of A In Each Trial. The Order Of Species D Which Is x Is Also Constant. Canceling Terms We

Have Left: $\text{Rate}_1 \text{ Rate}_2 = K[B]_1 Y K[B]_2 Y$ Equation 6 Consider The Following
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Very Rapidly. Also, The More Bonds Between Reacting Atoms In A Molecule, The
Slower The Reaction Rate. Why? More Energy Is Required To Separate The Molecule
Into Its “bits”. Substances With Strong Bonds (larger Bond Energies) Will React
Much More Slowly. Jan 3th, 2024

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So, To Account For Effective Collisions, Another Factor, P Called The Probability Or

Steric Factor Is Introduced. So, Rate = $PZ AB E-E A /RT$ The Activated Complex Theory Or Transition State Theory Reactants \rightleftharpoons Activated Complex \rightarrow Products This Theory Is Based On The Fact That Jan 6th, 2024

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