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4 Arrangement Of Electrons In Atoms1 S22s 2p63 23p64s1 C. Gallium 1s 22s 22p63s 3p63d104s 4p1 D. Copper 1s 22s 2p 63s23p 3d104s1 PROBLEMS Write The Answer On The Line To The Left. Show All Your Work In The Space Provided. 9. 1 1012 M What Is The Wavelength Of Light That Has A Frequency Of  $3 \times 10^4$  Hz In A Vacuum? 10. 3.3  $10^{19}$  18th, 2024Chapter 4 Arrangement Of Electrons In Atoms

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Quantum Number 1. Indicates The Fundamental Spin States Of An Electron In An Orbital 2. Two Possible Values For Spin,  $+1/2$ ,  $-1/2$  3. A Single Orbital Can Contain Only Two Electrons, Which Must Have 23th, 2024.

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Four Orbitals Are Pictured Below. The S-orbital Is Spherical In Shape. The P-orbital Is A 3-dimensional Dumb-bell Shape. Its 3 Orbitals Exist In Different Planes (called X, Y And Z) And Are Right Angles To One Another. The D-orbital Has Four Lobes, While The F

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