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Of Solution  $\chi_{\text{Sugar}} = (0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$  Mole Fraction Water Of  
Solution  $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$  (or  $1 - \chi_{\text{Sugar}}$ ) Molality  $M =$   
 $(.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$  Molarity  $M = (.292 \text{ Mol}) / (1 \text{ Apr 7th},$   
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K Is The Molal Freezing Point Depression And M Molality Is ... As  $1 \text{ atm} = 1.013 \times$   
 $10^5 \text{ Pa}$ , The Osmotic Pressure Is  $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$  Hence C

$$= \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$$
 As  $1 \text{ M}^3 = 1000 \text{ L}$ , This Corresponds To  $C = 1.5 \times 10^{-4} \text{ Mol L}^{-1}$

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