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Oxidation Reduction Examples With Answers

A.  $\text{ClO}_3^-$  -  $\text{ClO}_2$  Reductive Agent

B.  $\text{SO}_4^{2-}$  -  $\text{S}^{2-}$  - Reducing Agent

C.  $\text{Mn}^{2+}$  -  $\text{MnO}_2$  Oxidizing Agent

D.  $\text{Zn}^{2+}$  -  $\text{Cl}_2$  Oxidant

Write Balanced Equations For The Following Redox Reactions:

A.  $2 \text{NaBr} + \text{Cl}_2 \rightarrow 2 \text{NaCl} + \text{Br}_2$

B.  $\text{Fe}_2\text{O}_3 + 3 \text{CO} \rightarrow 2 \text{Fe} + 3 \text{CO}_2$  In Sour Solution

C.  $5 \text{CO} + \text{I}_2\text{O}_5 \rightarrow 5 \text{CO}_2 + \text{I}_2$  In The Basi

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